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Facile electrochemical synthesis of antimicrobial TiO₂ nanotube arrays

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Abstract: Infection-related complications have been a critical issue for the application of titanium orthopedic implants. The use of Ag nanoparticles offers a potential approach to incorporate antimicrobial properties into the titanium implants. In this work, a novel and simple method was developed for synthesis of Ag (II) oxide deposited TiO₂ nanotubes (TiNTs) using electrochemical anodization followed by Ag electroplating processes in the same electrolyte. The quantities of AgO nanoparticles deposited in TiNT were controlled by selecting different electroplating times and voltages. It was shown that AgO nanoparticles were crystalline and distributed throughout the length of the nanotubes. Inductively coupled plasma mass spectrometry tests showed that the quantities of released Ag were less than 7 mg/L after 30 days at 37°C. Antimicrobial assay results show that the AgO-deposited TiNTs can effectively kill the *Escherichia coli* bacteria. Although the AgO-deposited TiNTs showed some cytotoxicity, it should be controllable by optimization of the electroplating parameters and incorporation of cell growth factor. The results of this study indicated that antimicrobial properties could be added to nanotextured medical implants through a simple and cost effective method.

Keywords: TiO₂ nanotube arrays, anodization, AgO nanoparticles, antimicrobial, cytotoxicity

Introduction

Currently, bacterial infection is one of the most challenging issues for biomedical implants. Ag nanoparticles have strong antimicrobial function and this has been extensively studied.^{1,2} The antimicrobial function is possibly contributed by four mechanisms.^{1,3} The first is the attachment of Ag nanoparticles (1–10 nm) on the cell membrane. The second is that Ag nanoparticles penetrate into the cell and interact with sulfur and phosphorus in the cells. Ag nanoparticles can also produce reactive oxygen species (ROS).³ The last is the release of Ag ions from Ag particles to destroy the DNA replication ability.² The release mechanism of Ag⁺ from Ag nanoparticles includes the oxidation of Ag nanoparticles and Ag⁺ release under acidic conditions, which is described in the following equation.^{4,5}



Another critical issue for Ag in medical implants is the cytotoxicity problem. It was found that relatively small Ag nanoparticles led to lower cell viability.⁶ This was related to the large release of Ag ions due to the high surface energy of smaller Ag nanoparticles. The release of Ag ions from Ag nanoparticles is associated with one or two oxidized monolayers on the surface of the nanoparticles.⁷ For biomedical implant

applications, smaller Ag particle size does not always lead to better performance. The cell cytotoxicity of Ag nanoparticles can be controlled by reducing the total quantity of small Ag nanoparticles or utilizing relatively large nanoparticles.

In recent years, researchers have shown much interest in TiO_2 nanotubes (TiNTs) due to their unique properties for bio-applications.^{8–18} TiNT structures are particularly interesting because of their large specific surface areas and tube-like structure, which may supply many more reactive sites for the reaction. Due to the antimicrobial function of Ag nanoparticles, depositing Ag on TiNTs may bring better performance for biomedical implants. Some studies investigated the antimicrobial and cytotoxic effects of sputtering-coated Ag nanoparticles on TiNTs.^{19,20} It was found that just 0.01 mg/cm² Ag on TiNTs could effectively kill the *Staphylococcus epidermidis* without significant damage to the cells.

Many fabrication methods have been developed to deposit Ag onto and in TiNTs, including pulsed current deposition,²¹ successive ionic layer adsorption and reaction,^{22,23} photochemical reduction,^{24–26} layer by layer assembly,^{27,28} chemical assembly,²⁹ and sputtering.^{19,20} Pulsed current deposition method can produce a uniform deposition of Ag nanoparticles, but this method needs a special pulsed current control system. Successive ionic layer adsorption and reaction, layer by layer and chemical assembly require several different chemical reagents and multiple steps. Sputtering can produce a very consistent distribution of Ag nanoparticles, but sputtering is a relatively time consuming and high cost operation. Photochemical reduction is a relatively simple method, but this method requires different electrolytes for anodization and Ag depositing.

In this study, a novel and cost effective method was developed to synthesize the AgO-deposited TiNT arrays. Anodization and AgO-depositing processes were performed in the same electrolyte containing AgF to supply both the fluoride ions for Ti anodization and Ag ions for synthesis of AgO nanoparticles by electroplating at constant voltages. Only one DC power supply was required for both the anodization and Ag-electroplating processes. In the present work, bacterial and cell cytotoxic tests were performed to evaluate the antimicrobial and cytotoxic properties of the AgO-deposited TiNTs.

Materials and methods

Specimen preparation

Pure Ti foil (Alfa, 0.25 mm thick, 99.5%) was selected as the substrate to form the TiNT layer by electrochemical anodization. Prior to anodization, the samples were polished with sand paper 300, 400 and 1,200 and then cleaned in ethanol for 10 minutes. During each anodization operation, the back side and edges of the samples were covered with nail polish leaving four 10×10 mm² areas uncovered. Figure 1 shows the setup for anodization and Ag-depositing (Ag electroplating). The anodization system was a two-electrode electrochemical cell. Ti foils served as the anode while a carbon rod was used as the cathode. All samples were anodized in fresh electrolyte containing 0.686 g AgF, 98 mL ethylene glycol and 2 mL deionized water. The anodization process was performed at 60 VDC for 90 minutes in a 25°C water bath. During anodization, TiNTs were formed at the Ti substrate. Ag^+ in the electrolyte was adsorbed onto the surface of the carbon cathode as the electric field supplied electrons for Ag

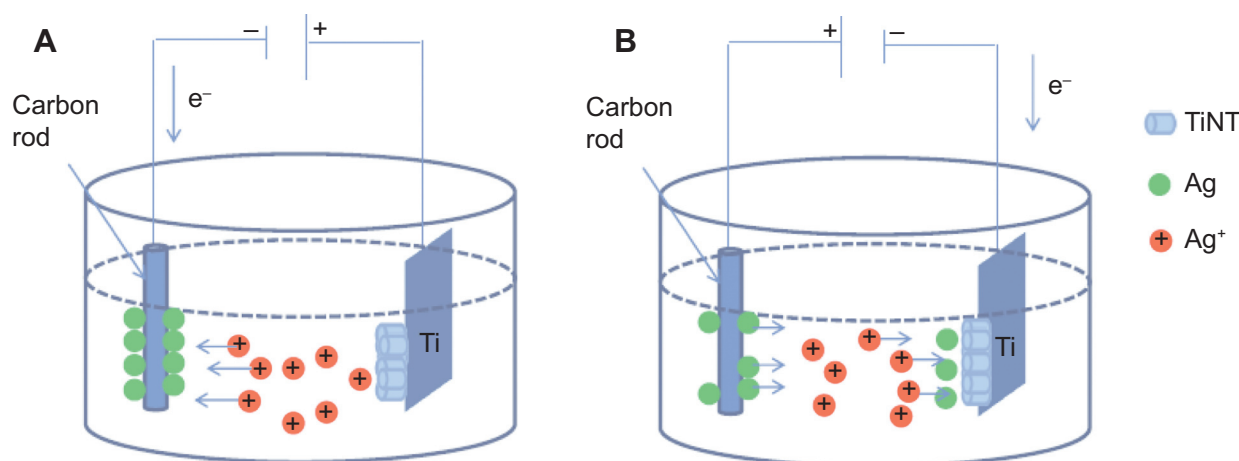


Figure 1 Schematic diagram of AgO-deposited TiNT formation.

Notes: (A) Anodization process is shown. Ag ions moved toward carbon cathode, which was negatively biased with respect to Ti. Ag atoms were deposited on the cathode while TiNT were formed on Ti foil which was positively biased. (B) Ag electroplating process is shown. The polarity was reversed and Ag^+ ions moved toward negatively biased Ti with respect to carbon cathode. Ag atoms deposited on the TiNT layer on Ti.

Abbreviation: TiNT, TiO_2 nanotubes.

deposition. For the Ag depositing process, the samples were first covered by Spectra/Por® 7 semi-permeable membrane (Spectrum Laboratories, Inc., Rancho Dominguez, CA, USA). Then the electrodes were switched and 20 VDC was applied to rapidly deposit Ag nanoparticles onto the TiNTs. During the Ag deposition process, Ag on the cathode lost electrons and dissolved into the electrolyte. Under the electric field, Ag⁺ in the electrolyte migrated into the TiNTs, took up electrons, and formed Ag nanoparticles. Three Ag-deposition times were applied, 30, 60, and 90 seconds. These samples were denoted as TiNT-Ag30s, TiNT-Ag60s, and TiNT-Ag90s, respectively. After Ag-deposition, the samples were rinsed in acetone, isopropanol, and deionized water each for 5 minutes.

Characterization and Ag release tests

The surface topography of AgO-deposited TiNTs was observed by Hitachi S-4700 scanning electron microscopy (Hitachi Ltd, Tokyo, Japan). The elemental composition of the AgO-deposited TiNTs were measured by energy-dispersive X-ray spectroscopy. The crystal structure of the AgO-deposited TiNTs was characterized by JEOL JEM-3010 (JEOL, Tokyo, Japan) transmission electron microscope (TEM). For Ag release tests, the samples were incubated in 5 mL α -Minimum Essential Medium (α -MEM), 10% fetal bovine serum (FBS), and 1% pen-strep at 37°C. The Ag⁺ concentration in the medium was analyzed by Perkin-Elmer® Optima 7000 DV (PerkinElmer Inc., Waltham, MA, USA) inductively coupled plasma mass spectrometry.

Cell culture

MC3T3 mouse osteoblast cells were selected for cytotoxicity and proliferation tests. The cells were cultured in a complete medium composed of α -MEM, 10% FBS, and 1% pen-strep and incubated in a humidified atmosphere containing 5% CO₂ at 37°C.

Cell proliferation tests

The MC3T3 cells were grown in 2 mL of medium with AgO-deposited TiNT samples (10×10 mm²) in the 12-well plate. Before the cell tests, all samples were sterilized with 70% ethanol and rinsed with sterile water. The DNA content in the medium was determined fluorometrically using PicoGreen assay kit (Thermo Fisher Scientific, Waltham, MA, USA). The medium was removed and the samples were rinsed with phosphate buffer saline (PBS) solution. Then, the cells were lysed using 1 mL proteinase K solution at 37°C for 1 hour. One hundred μ L proteinase K solution was placed in triplicate in a 96-well plate and mixed

with 100 μ L PicoGreen. The plate was incubated at room temperature for 10 minutes in the dark and then read on a Fluoroskan Ascent™ FL fluorescent plate reader (Thermo Fisher Scientific). For each data point, four samples were tested. The statistical significance was calculated by analysis of variance (ANOVA) test.

Lactate dehydrogenase activity assay

The cytotoxicity of Ag-deposited TiNTs was determined by lactate dehydrogenase (LDH) assay. After 1 and 4 days, the supernatant culture media was collected and used for the LDH assay. The assay was operated based on the manufacturer's instructions (Promega Corporation, Fitchburg, WI, USA). Similar to the cell proliferation tests, four samples were tested for each data point. The statistical significance was calculated by ANOVA test.

Antimicrobial tests

Antimicrobial effects of AgO-deposited TiNT samples were evaluated using *E. coli* (One Shot® Mach1™ T1 Phage *E. coli* from Thermo Fisher Scientific) as previously reported.³⁰ Briefly, *E. coli* was initially cultured in 15 mL Luria broth (Sigma-Aldrich Co., St Louis, MO, USA, L3522) overnight at 37°C and plated on agar plates at a density of approximately 10⁷ cfu. Ti-nanotubes were cut into 10×10 mm² squares, sterilized with ethanol, and rinsed with sterile water before placing them on top of the inoculated agar plates. The agar plates were incubated at 37°C for 24 hours (hr). The zone of inhibition (ZOI; clear zone formed around the samples) was measured using a ruler and values reported are average of three independent samples for each condition.

The viability of bacteria on the surface of AgO-deposited TiNTs was measured based on differential fluorescence staining for live and dead bacteria as reported previously.³¹ Briefly, *E. coli* from a glycerol stock was inoculated into 15 mL of Luria broth and incubated at 37°C overnight. Approximately 8×10⁶ cells from the culture were inoculated on the surface of nanotubes with fresh media that was incubated overnight at 37°C. All samples incubated with bacterial culture overnight were replaced with fresh media and were further incubated at 37°C for 7 days with fresh media being replaced every 24 hr. This was followed by adding fresh bacterial suspension to nanotubes and incubating it for another 7 hr before rinsing with PBS. Samples were stained using acridine orange (4 μ g/mL) and ethidium bromide (0.5 μ g/mL) for 15 minutes in the dark and observed under fluorescence microscopy (EVOS-fl, AMAG Pharmaceuticals Inc., Lexington, MA, USA). The

images were taken with red fluorescent protein (RFP) filter for ethidium bromide and green fluorescent protein (GFP) filter for acridine orange. All the images were acquired using 60× objective. The live cells stained with acridine orange appear green whereas the dead cells stained with ethidium bromide appear red.

Results

Characterization of AgO-deposited TiNTs

Figure 2A–C shows the surfaces of AgO-deposited TiNTs for different electroplating times. Nanoparticles (about 80 nm) were found on the surface of the TiNTs. With increased deposition time, more nanoparticles were observed. Energy-dispersive X-ray spectroscopy (Figure 2D) shows a small peak of Ag at 3 keV. Figure 2E and F show the topography and surface potential of TiNT-Ag60s by atomic force microscopy (AFM). The size of surface nanoparticles and the TiNT diameters were similar. So, it was difficult to distinguish the surface nanoparticles from TiNT substrate by AFM topography image (Figure 2E). Due to the potential difference between AgO and TiO₂, AgO nanoparticles showed some much brighter areas in Figure 2F, which means a higher surface potential is associated with AgO. To confirm the crystal structure of the nanoparticles and TiNT substrate, TEM images and diffraction patterns were obtained. Figure 3 shows the TEM images of the TiNT-Ag60s. Based on the diffraction pattern in Figure 3A, the TiNT substrate was amorphous. As shown in Figure 3A and B, about 10 nm black nanoparticles attached inside the TiNT. Based on the scanning electron microscopy and TEM images, the Ag nanoparticles fabricated in present work were composed of relatively large nanoparticles (about 80 nm) at the surface and small nanoparticles (about 10 nm) inside the TiNTs. Figure 3C shows the atomic resolution images of the black particles. Based on the diffraction pattern (Figure 3D), the black nanoparticles are AgO. Figure 4 shows the high resolution X-ray photon spectroscopy (XPS) spectrum of Ag 3d in TiNT-Ag60s. The binding energy peaks of Ag 3d are at 367.4 eV and 373.4 eV. Ag 3d_{5/2} peak for AgO is at 367.4 eV.³² So, the XPS spectrum of Ag 3d in TiNT-Ag60s (Figure 4) further confirms that the nanoparticles are AgO. Both TEM and XPS results clearly show that AgO nanoparticles were on the TiNTs. Based on the Ag–H₂O Pourbaix diagram,³³ Ag particles should have formed during the electroplating process. But the electrolyte was acidic, which was more favorable for the oxidation of Ag. It was also found that AgO tended to form during

electrochemical oxidation in acid and neutral solution of Ag⁺ salt.^{34,35} In another aspect, the small Ag nanoparticles had a high surface energy which made them more susceptible to oxidation in an acidic solution. Both the acidic environment and the high surface energy might have caused the transition from Ag to AgO.

Ag-released test results

Figure 5 shows Ag release from the TiNTs as a function of time in the cell culture medium at 37°C. The total Ag released from TiNT control samples was less than 2 mg/L after 30 days. The released Ag ions from the TiNT control samples were from the residual electrolyte in the TiNTs. The highest cumulative Ag release was less than 7 mg/L after 30 days. The amount of Ag released from the AgO-deposited samples increased with the electroplating time. The basic Ag release mechanism has been described by equation 1 and 2. Ag nanoparticles firstly interact with oxygen to form Ag₂O and then Ag₂O reacts with hydrogen ions to form Ag⁺ and water.^{4,5} In our study, AgO may have followed the similar reaction as equation 2 to form Ag⁺. With longer Ag depositing time, more Ag was deposited onto the TiNTs. More Ag nanoparticles supplied more surface area for the release of Ag ions. The amount of Ag released from Ag nanoparticles followed an exponential curve, which is described by:³⁶

$$[Ag^+]_{\text{released}} = [Ag]_{\text{initial}} [1 - \exp(-at)] \quad (3)$$

Where $[Ag]_{\text{initial}}$ is the initial Ag concentration, a is fitting parameter and t is time. In the current work, the release rate of Ag⁺ decreased in longer release time, which was consistent with the exponential curve.

Cell cytotoxicity test

The cytotoxicity of the substrates on MC3T3 osteoblasts was evaluated by LDH assay. As shown in Figure 6A, on Day 1 all the samples show similar LDH activity ($P > 0.05$). After 4 days, all AgO-deposited samples show significantly higher LDH activity than pure Ti samples ($P < 0.01$). Even the TiNT samples showed a higher LDH activity than the Ti specimen. The cytotoxicity caused by TiNT samples is due to the Ag ions trapped in the TiNT during the anodization process. On Day 4, the LDH activities of TiNT-Ag30s and TiNT-Ag60s were a little higher than the TiNT samples. But the difference is statistically insignificant. Only TiNT-Ag90s showed significantly higher LDH activity than TiNT samples.

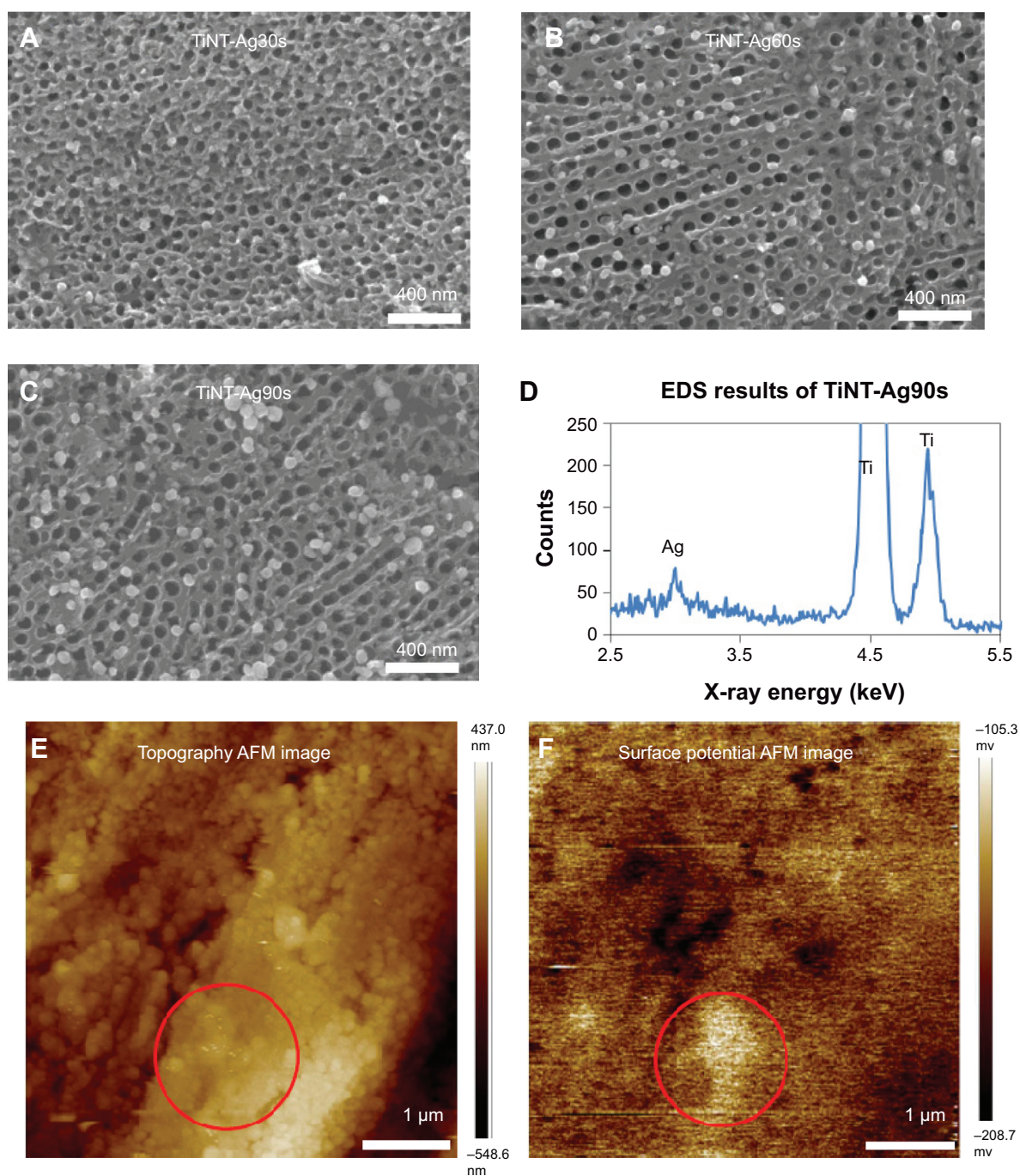


Figure 2 SEM images of AgO-deposited TiNT.

Notes: Treated with different Ag electroplating time: (A) 30s (TiNT-Ag30s), (B) 60s (TiNT-Ag60s), (C) 90s (TiNT-Ag90s) and (D) EDS results of TiNT-Ag90s. AFM images of TiNT-Ag60s: (E) topography and (F) surface potential. The circular marked area in the topography image is the area with high surface potential. SEM images show nanoparticles (about 80 nm) on the surface of the Ag-deposited TiNT surface. More nanoparticles were formed on samples with longer electroplating time. The EDS data show a small peak of Ag at 3 keV. AFM surface potential image shows some bright areas related to the higher Ag potential compared to TiNT substrate.

Abbreviations: SEM, scanning electron microscopy; TiNT, TiO₂ nanotubes; EDS, energy-dispersive X-ray spectroscopy; AFM, atomic force microscopy.

Cell proliferation

The amount of DNA released from lysed MC3T3 cells grown on different substrates is shown in Figure 6B. On days 1 and 4, the cell growth on the TiNT samples, as represented by the amount of lysed DNA, was about 50% less than that of Ti

control samples. The AgO-deposited samples showed much lower cell proliferation than the Ti control samples. The difference among Ag30s, Ag60s, and Ag90s samples was not significant after 1 day. But on day 4, the cell numbers decreased with the increased Ag electroplating time.

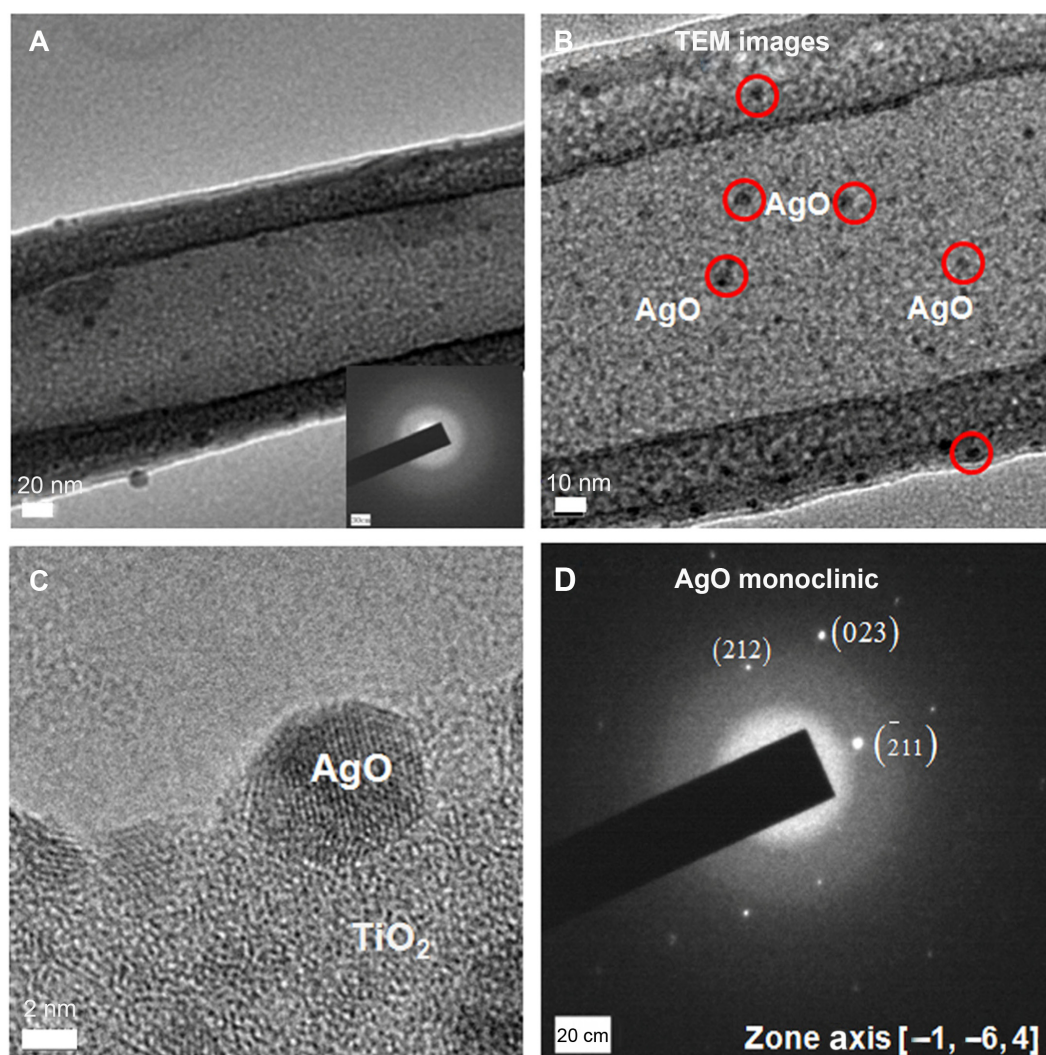


Figure 3 TEM images of TiNT-Ag60s.

Notes: (A) Morphology of AgO-deposited TiNT, (B) enlargement of black AgO nanoparticles on TiNT (red circles indicate some black AgO particles). (C) High resolution image of AgO nanoparticles and (D) diffraction pattern of AgO. Some 5–10 nm black particles were found inside the TiNT. Based on diffraction pattern, the crystal structures of black nanoparticles and substrate were monoclinic AgO and amorphous TiO_2 , respectively.

Abbreviations: TiNT, TiO_2 nanotubes; TiNT-Ag60s, TiNT treated with Ag electroplating time of 60 seconds; TEM, transmission electron microscope.

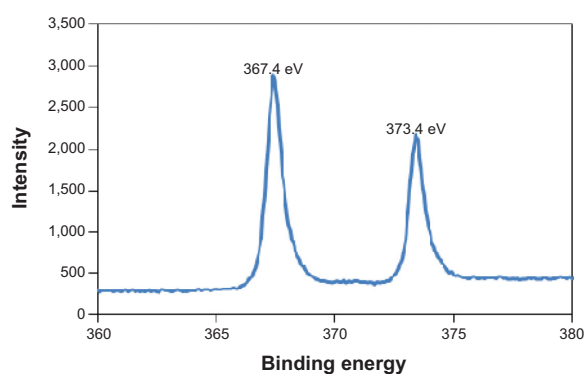


Figure 4 High resolution X-ray photon spectroscopy (XPS) spectrum of Ag 3d in TiNT-Ag60s (367.4 eV is the Ag $3d_{5/2}$ peak for AgO).

Note: The XPS spectrum of Ag 3d confirmed that the nanoparticles were composed of AgO.

Abbreviation: TiNT, TiO_2 nanotubes; TiNT-Ag60s, TiNT treated with Ag electroplating time of 60 seconds.

The current work showed cytotoxicity for MC3T3 cells apparently due to the relatively high Ag release from the AgO-deposited TiNT samples.

Bacterial test

Figure 7 shows the inhibition zone test results by control Ti, TiNTs and AgO-deposited TiNTs on *E. coli* bacteria. Ti and TiNT samples did not show any significant inhibition zone, even though it was shown that the TiNT samples released a small amount of Ag^+ . All of the AgO-deposited TiNT samples showed a clear inhibition zone, which indicated antimicrobial effects of these samples. The diameter of the bacteria inhibition zone was similar for all the AgO-deposited samples. Based on the Ag release in Figure 5, the amount of Ag was

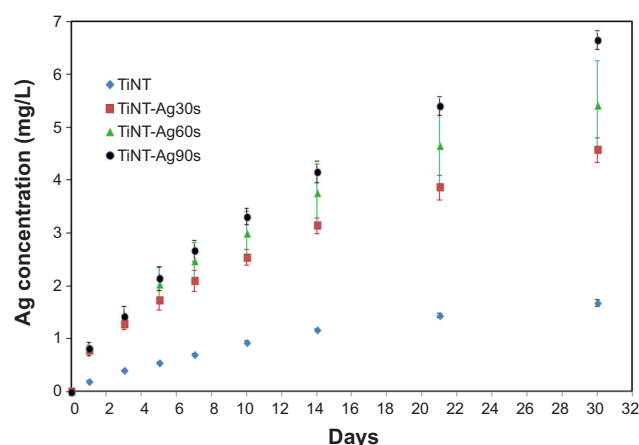


Figure 5 Cumulative release of Ag from TiNT and AgO-deposited samples soaked in cell culture medium at 37°C.

Notes: Experiment results were expressed as means \pm standard deviation of the means of the samples ($n=3$). The released Ag was from the trapped electrolyte in the TiO₂ nanotubes. The amount of Ag released from the AgO-deposited samples increased with the electroplating time. The Ag release followed the exponential curve. Ag30s, 60s, or 90s refer to treatment with an Ag electroplating time of 30, 60, or 90 seconds, respectively.

Abbreviation: TiNT, TiO₂ nanotubes.

about 0.8 mg/L for all AgO-deposited samples on Day 1. Only after 7 days, did samples with longer Ag electroplating times release more Ag into the medium. Although, Ag is a leaching antimicrobial agent, 1 day is not long enough to show the effect of different Ag depositing times. This explains why all the AgO-deposited samples showed similar bacteria inhibition zones.

The antimicrobial effect of AgO nanoparticles was also confirmed by bacteria viability after 7 days incubation. As shown in Figure 8A, large amounts of viable bacteria were

found at the surface of the Ti plate (Figure 8B). There were relatively small amounts of live bacteria on the TiNT samples. Due to the residual Ag⁺ from the electrolyte, the fluorescent images of the TiNT samples show relatively few living cells (Figure 8C) and some dead bacteria (Figure 8D). With the increase of Ag electroplating time, less bacteria were found on AgO-deposited samples. No live bacteria were found on TiNT-Ag60s and TiNT-Ag90s samples. In the present study, the TEM analysis (Figure 3A) indicated that the TiNT substrate had an amorphous structure. The antimicrobial effects of amorphous TiNT is negligible.³⁷ Therefore, the strong antimicrobial effects of AgO-deposited TiNT samples should be attributed to the AgO nanoparticles.

Discussion

In this study, the antimicrobial effects of AgO-deposited TiNTs were shown to be caused by Ag⁺ release from AgO nanoparticles. Xiu et al⁵ compared the antimicrobial properties of Ag nanoparticles under aerobic and anaerobic environments. It was found that the quantities of Ag⁺ released from Ag nanoparticles decreased greatly in an anaerobic environment. Correspondingly, Ag nanoparticles in the anaerobic environment did not induce significant death of *E. coli*. The antimicrobial behavior comparison of Ag nanoparticles and Ag⁺ also showed that the bacteria inhibition zones were mainly determined by Ag⁺ concentration for both Ag nanoparticles and AgNO₃ solution.³⁸ Ag ions can interrupt bacteria cellular processes and produce ROS.³⁹ These ROS can increase the permeability of the bacteria membrane, which make the bacteria more susceptible to the antimicrobial

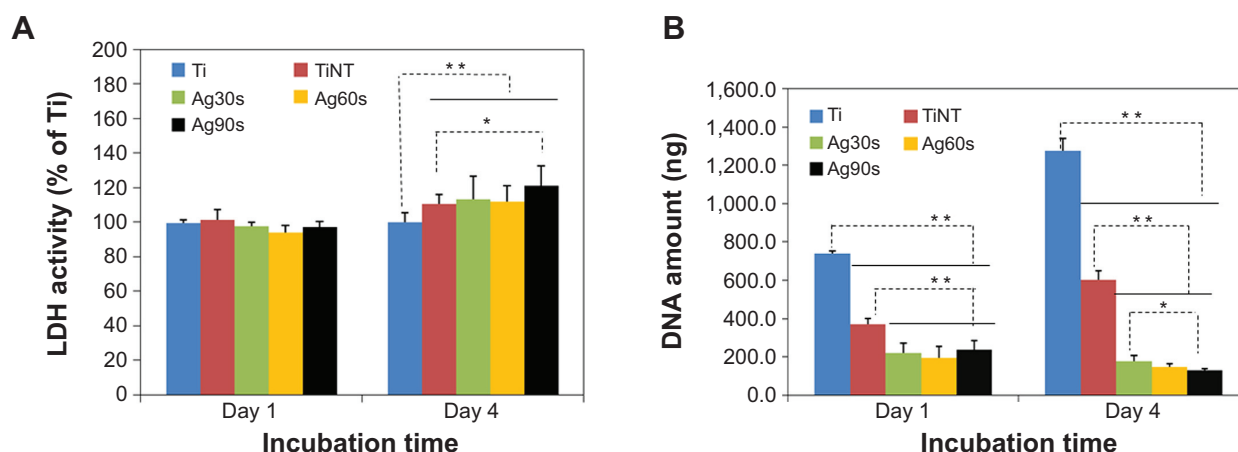


Figure 6 Cellular activity analysis of TiO₂ nanotubes.

Notes: (A) Cell viability test results by LDH assay for 1 and 4 days on the samples ($n=4$). At Day 1, the LDH activity of all the samples were similar. After 4 days, the TiNT and AgO-deposited samples showed a higher LDH activity, which means increased cytotoxicity to MC3T3 cells. (B) Cell proliferation determined by DNA amount after MC3T3 cells were cultured for 1 and 4 days on the samples ($n=4$). The DNA amount of TiNT was about half of Ti samples on Day 1 and Day 4, which indicated that cell proliferation of TiNT samples is about 50% less than that of Ti samples. AgO-deposited samples showed much lower cell proliferation than that of Ti samples on Day 1 and 4. AgO nanoparticles induced cytotoxicity for MC3T3 cells at relatively low Ag release. The statistical significance was calculated by ANOVA test. (** $P<0.01$ and * $P<0.05$). Ag30s, 60s, or 90s refer to treatment with an Ag electroplating time of 30, 60, or 90 seconds, respectively.

Abbreviations: ANOVA, analysis of variance; LDH, lactate dehydrogenase; TiNT, TiO₂ nanotubes.

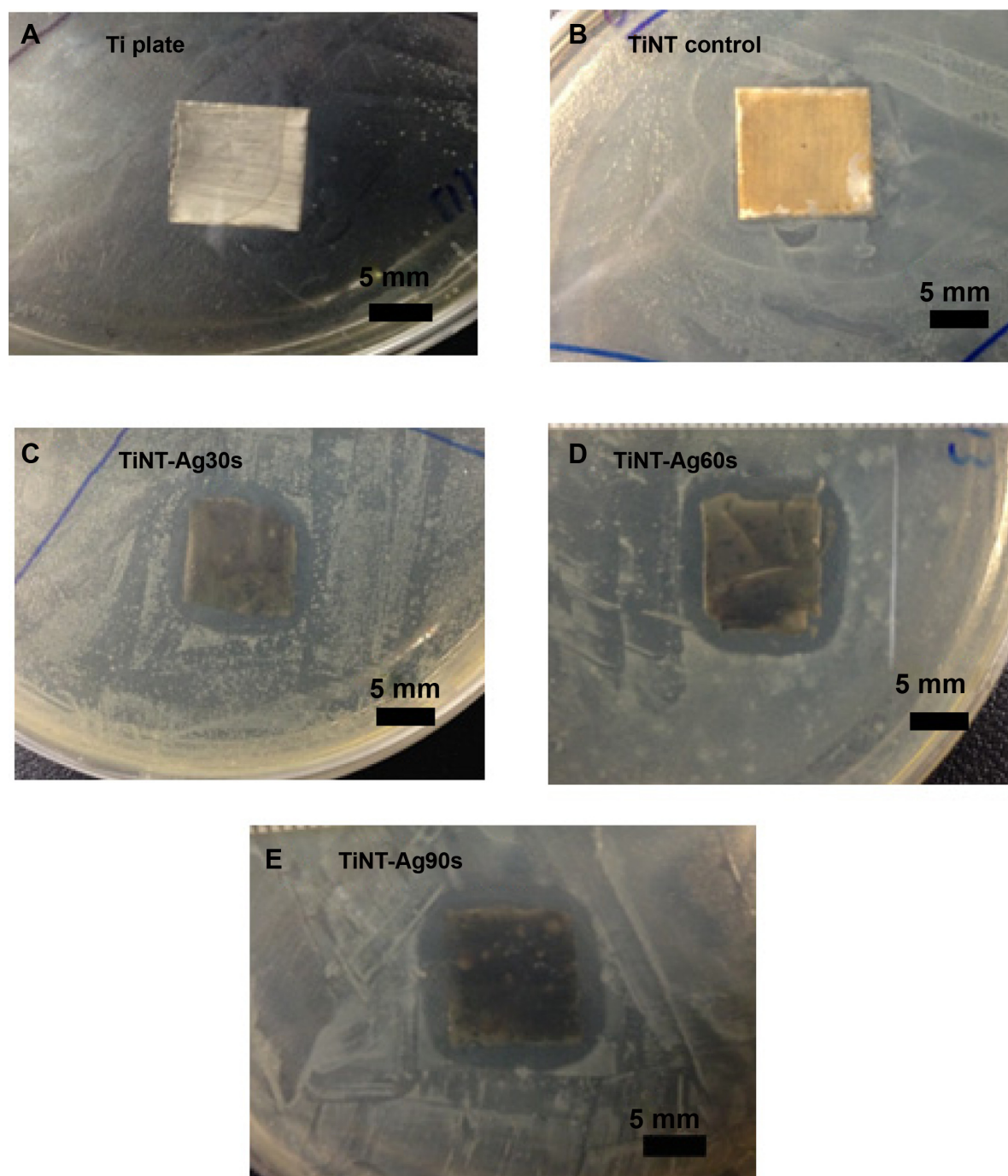


Figure 7 Zone of inhibition (Zol) for samples exposed to *Escherichia coli* on Day 1.

Notes: Ti and TiNT samples did not show Zol. All AgO-deposited samples showed similar Zol (diameter: 7 mm) due to similar Ag release on Day 1. AgO-deposited TiNT showed good antimicrobial effects. (A) Ti plate, (B) TiNT control, (C) TiNT-Ag30s, (D) TiNT-Ag60s, (E) TiNT-Ag90s. Ag30s, 60s, or 90s refer to treatment with an Ag electroplating time of 30, 60, or 90 seconds, respectively.

Abbreviation: TiNT, TiO₂ nanotubes.

agent. The antimicrobial effects of AgO-deposited TiNTs should be mainly attributed to the Ag⁺ released from AgO nanoparticles. One should note that the species most often associated with infections in orthopedic devices are *Staphylococcus aureus* and *S. epidermidis*. Others such as: *Acinetobacter baumannii* and *E. coli* are not predominant.

Further bacteria testing is needed to better understand the reaction of *S. aureus* and *S. epidermidis* to Ag release from our nanotubes.

As shown in Figure 6, AgO-deposited TiNT samples showed cytotoxicity to the MC3T3 osteoblast cells. Even TiNT samples, which only contained about 0.5 mg/L Ag⁺ on Day 4,

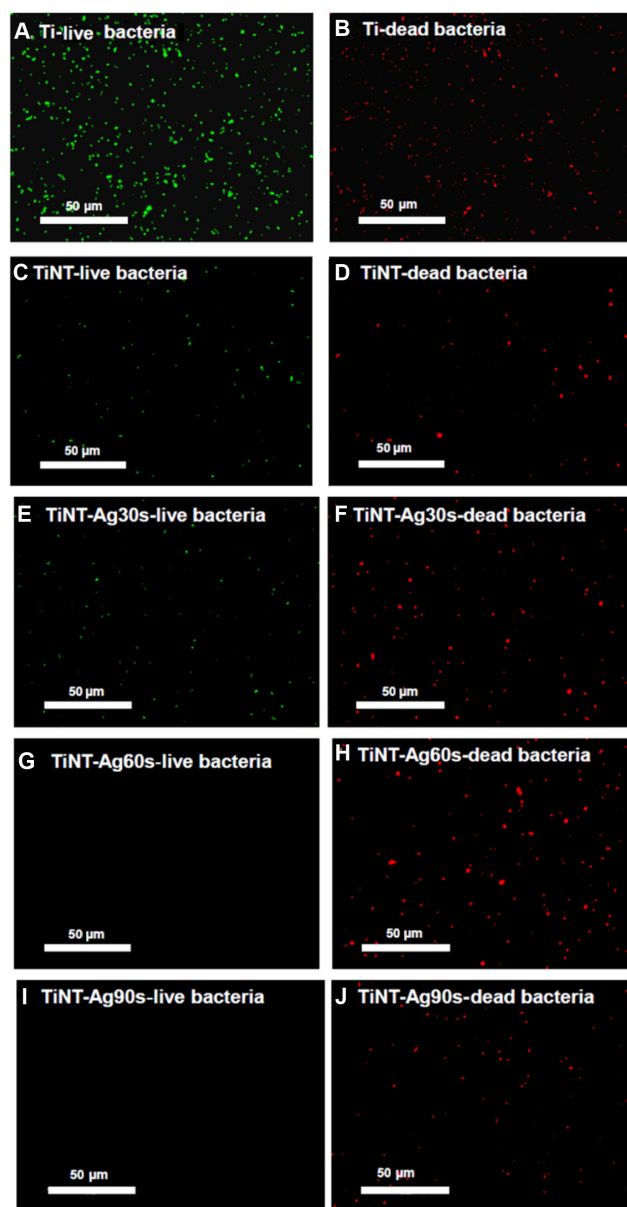


Figure 8 Fluorescent images of bacteria viability after 7 days of incubation (acridine orange and ethidium bromide fluorescence staining).

Notes: Live bacteria appear green and dead cells are red. A lot of bacteria were found at the Ti plate samples. TiNT samples' surface had some dead bacteria. The surface of TiNT-Ag60s and Ag90s samples did not show any viable bacteria. Ag30s, 60s, or 90s refer to treatment with an Ag electroplating time of 30, 60, or 90 seconds, respectively.

Abbreviation: TiNT, TiO₂ nanotubes.

showed a 50% reduction of the cell numbers. The cytotoxicity test in the current work demonstrated that MC3T3 cells are sensitive to the Ag⁺ even at relatively low concentration (~0.5 mg/L). This is in agreement with the *in vitro* study about Ag-deposited TiNTs where a similarly low Ag⁺ cytotoxicity concentration was reported.³¹ In general, it has been suggested that Ag ions are considered to be biocompatible only at relatively low concentrations. For osteoblasts from calvariae of

1–2 day old mice, the cytotoxic concentration of Ag ions was found to be ~5 mg/L.³⁸ For osteoblast cells of a neonatal rat, 2 mg/L Ag⁺ was still safe.³⁰ The Ag⁺ cytotoxic concentration for human mesenchymal stem cells and osteoblast cells are about 10 mg/L.⁴⁰ The present work showed cytotoxic behavior of Ag ions at a lower concentration. While it is not fully clear, several scenarios may contribute to this discrepancy. First, it is known that the use of different cell lines may explain this difference. Second, the cytotoxic concentration greatly depends on the original number of cells. Generally, higher initial cell numbers required more Ag⁺ to cause cell cytotoxicity. In the present study, the initial seeded cells were exposed to AgO-deposited samples even at the beginning, therefore affecting the first stage of proliferation and differentiation. The true cytotoxicity behavior of our AgO-deposited TiNTs should be studied *in vivo*. In fact, the *in vivo* tests of some Ag-TiNTs showed a strong antimicrobial effect without serious cytotoxicity for osteoblast cells even though the *in vitro* test suggested cytotoxicity of their Ag-deposited TiNTs.^{31,41} This is possibly related to the large quantities of osteoblast cells available near the implant site. Another scenario would be that Ag nanoparticles could have induced the bone growth by enhancing the mesenchymal stem cells' differentiation to osteoblast and/or have induced bone remodeling by enhanced osteoblast-osteoclast activity. In *in vivo* models many factors are involved in bone growth and osseointegration for which the *in vitro* tests are not sufficient and representative enough to form a conclusive argument on cytotoxicity of Ag. Therefore, it is important to confirm the cytotoxic concentration threshold by *in vivo* tests in future research.

Ag release from AgO nanoparticles was determined by many factors, especially the particle size. With the reduction of particle size, more surface area is available for AgO to react with water to produce Ag⁺. Smaller Ag nanoparticles can facilitate fast dissolution and release of Ag⁺.^{30,31,38,42} In the present study, two different AgO nanoparticles were fabricated, including 80 nm at the surface and 5–10 nm inside the nanotubes. The small AgO nanoparticles inside the TiNTs could possibly release Ag⁺ faster and may prevent the patient from infection at an early stage. The large AgO nanoparticles at the surface may supply long-lasting antimicrobial effects, which is a critical issue for biomedical implants. The AgO-deposited TiNTs prepared by electroplating could possibly supply long-term antimicrobial effects. Further experimental work on the release of Ag⁺ from large AgO particles versus smaller particles would be needed to clarify this conclusion.

To avoid *in vitro* cell cytotoxicity in the AgO-deposited TiNTs, it is necessary to reduce Ag release. It has been

found that 0.2 mg/L Ag⁺ concentration could effectively kill bacteria.⁵ It is possible that the bacteria can be killed without inducing cell cytotoxicity at relatively low Ag⁺ concentration. There may be several methods to reduce the cytotoxicity of AgO-deposited TiNTs. First, the Ag release can be reduced by modifying the Ag electroplating parameters. Another method is to add cell growth factor into the AgO-deposited TiNTs. For example, bone morphogenetic protein-2 (BMP-2) may be coated on TiNTs to improve the growth of osteoblast cells.^{43–45} More cell growth can increase the cytotoxicity threshold concentration and may reduce cell cytotoxicity induced by Ag ions. Future work will focus on the optimization of the Ag-deposition parameters and incorporation of BMP-2 or other growth factor into the AgO-deposited TiNTs. In vivo testing is necessary to confirm the Ag⁺ cytotoxic concentration for practical applications of AgO-deposited TiNTs as biomedical implants.

Conclusion

In this study, a novel and simple method was developed to deposit AgO nanoparticles on and in TiNTs by electroplating at constant voltages. The AgO nanoparticles in the TiNTs were mainly composed of 80 nm diameter at the surface and 5–10 nm inside the TiNTs. Bacteria ZoI and viability tests indicated that AgO-deposited TiNTs effectively killed *E. coli* bacteria. The small AgO nanoparticles can release Ag⁺ quickly and possibly avoid infection at the initial stage, whereas the larger nanoparticles may supply long-lasting antimicrobial effects. Current work found that the Ag⁺ cytotoxicity concentration of MC3T3 cells is low and this is related to the initial cell numbers. Although the AgO-deposited TiNTs showed some cell cytotoxicity, this may be reduced by optimization of the Ag electroplating parameters and incorporation of a growth factor. The results of this study indicate that antimicrobial properties can be added to nanotextured medical implants through a simple and cost effective method.

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Disclosure

The authors report no conflicts of interest in this work.

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