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## Electrochemical glucose biosensor based on nickel oxide nanoparticle-modified carbon paste electrode

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### Abstract

In the present work, we designed an amperometric glucose biosensor based on nickel oxide nanoparticles (NiONPs)-modified carbon paste electrode. The biosensor was prepared by incorporation of glucose oxidase and NiONPs into a carbon paste matrix. It showed good analytical performances such as high sensitivity ( $367 \mu\text{A mmolL}^{-1}$ ) and a wide linear response from  $1.9 \times 10^{-3} \text{ mmolL}^{-1}$  to  $15.0 \text{ mmolL}^{-1}$  with a limit of detection ( $0.11 \mu\text{molL}^{-1}$ ). The biosensor was used for the determination of glucose in human serum samples. The results illustrate that NiONPs have enormous potential in the construction of biosensor for determination of glucose.

**Keywords:** amperometric glucose biosensor, carbon paste electrode, glucose oxidase, nickel oxide nanoparticles

### Introduction

Diabetes mellitus is a group of carbohydrate metabolic disorders affecting about 150 million people around the world. This disease leads to death and failure of several organs such as heart, kidneys, and eyes (Mercado and Moussy 1998, Moussy et al. 1994, Privett et al. 2008). This makes glucose biosensing very important because accurate and rapid determination of blood glucose is crucial for treatment and management of diabetic patients.

Conventional techniques for glucose determination such as chemiluminescence (Zhu et al. 2002), colorimetric (Wang et al. 2009), and capillary electrophoresis (Wang et al. 2011) are complex, expensive, and time consuming. In comparison, enzyme-based amperometric biosensors have advantages over nonenzymatic ones; such as the ease of operation, low cost of fabrication, high selectivity, and operational stability for real-time measurements (Guo et al. 2011).

In amperometric glucose biosensors, glucose oxidase (GOD) catalyzes the oxidation of glucose to gluconolactone and oxygen reduces to hydrogen peroxide ( $\text{H}_2\text{O}_2$ ). An enzyme electrode for glucose determination is based on the

detection of enzymatically generated  $\text{H}_2\text{O}_2$  during the enzymatic reaction (Yang et al. 2009, Usman Ali et al. 2010).



The  $\text{H}_2\text{O}_2$  can be oxidized or reduced directly at ordinary solid electrodes. However, these processes in analytical applications are limited by slow electrode kinetics and high overpotential which may incur large interferences from other existing electroactive species available in real samples such as ascorbic acid, uric acid, urea, paracetamol, *etc.* Thus, the determination of glucose is mainly focused on the electrode modifications in order to decrease the overpotential and increase the electron transfer kinetics. For this reason, a large range of materials such as redox mediators (Jiang et al. 2008), conducting polymers (Çolak et al. 2012), carbon nanotubes (Zargoosh et al. 2012) have been employed to construct of the electrochemical sensors for detection of glucose. On the other hand, in past decades, nanomaterials have attracted great research interest owing to their desirable chemical, physical, and electronic properties that are different from those of bulk materials.

Recently, NiONPs are considered as important metal oxides with attractive electrochemical, thermal, optical, and mechanical properties with their high surface to volume ratio (Lata et al. 2012, Hotovy et al. 2002). In constructing enzyme-based biosensors, NiONPs are favorable because of their high isoelectric point for physical adsorption of enzyme molecules by electrostatic interactions (Li et al. 2008). The electrochemical applications of NiONPs in glucose (Ding et al. 2011, Mu et al. 2011), myoglobin (Zong et al. 2010), DNA (Noorbakhsh and Salimi 2011), and  $\text{H}_2\text{O}_2$  (Salimi et al. 2007a) sensors have been applied in various fields.

In the present study, we developed a new glucose biosensor based on NiONPs-modified carbon paste GOD electrode for amperometric detection of glucose. The optimum working conditions such as working potential, enzyme unit,

pH, and temperature were determined. The analytical performance factors and application of the developed glucose biosensor to real samples were investigated.

## Experimental

### Chemicals and reagents

GOD (EC 1.1.3.4, from *Aspergillus niger*, EC 1.1.3.4 39680 Units/g solid) and nickel oxide nanoparticles (<50 nm) were purchased from Sigma. Glucose, graphite powder, and paraffin oil were obtained from Fluka. Uric acid (UA), ascorbic acid, creatinine, urea, glucose, glutaraldehyde, and bovine serum albumine (BSA) were also purchased from Sigma. H<sub>2</sub>O<sub>2</sub> (35%, v/v aqueous solution), di-sodium monohydrogen phosphate heptahydrate, and sodium dihydrogen phosphate dihydrate were acquired from Riedel de Haën. All other chemicals were of analytical grade and used without further purification. Glucose stock solutions were allowed to mutarotate for at least 24 h at room temperature prior to use and stored at 4°C. All solutions were prepared with ultrapure water.

### Instrumentations

All electrochemical experiments were performed with a computer controlled on IVIUM-STAT electrochemical analyzer (Ivium Technologies, The Netherlands) and a three-electrode cell stand (Bioanalytical Systems, Inc., USA). A conventional three-electrode system comprises with a modified carbon paste electrode as the working electrode, a Pt wire (BAS MW 1034) as the counter electrode, and an Ag/AgCl including 3 mol L<sup>-1</sup> NaCl (BAS MF 2052) as the reference electrode. Temperature control was accomplished with a Grant LTD GG thermostat. The pH values of the buffer solutions were measured with an ORION model 720A pH-ion meter. The ultrapure water (18.2 MΩ cm) was obtained from Elga Purelab device (Veolia Water Systems Ltd., UK).

### Preparation of unmodified carbon paste and modified carbon paste electrodes

For unmodified carbon paste electrode (UCPE) construction, carbon paste was prepared by mechanically hand mixing graphite powder (70.0% w/w) and paraffin oil (30.0% w/w) until homogeneous paste was obtained. NiONPs-modified CPE (NiONPs-CPE) was prepared by mixing NiONPs with graphite powder in a ratio of 1:3 (w/w), followed by the incorporation of paraffin oil, and mixed regularly until a uniform paste was obtained. For the GOD enzyme and NiONPs incorporated carbon-paste electrode (GOD-NiONPs-CPE), the enzyme solution containing various units of GOD, 10 μL of glutaraldehyde (1.25% (v/v)) and 1.0 mg BSA were added into the graphite powder-NiONPs mixture. Finally, carbon paste was obtained by mixing with 15 μL paraffin oil for 20 min. A portion of the given paste was placed firmly into a cavity of polyetheretherketone (PEEK) rod electrode bodies (with 3 mm internal diameter) and the surfaces of electrodes were smoothed on a weighing paper before starting new experiment. The obtained electrodes were stored in the refrigerator at +4°C when not in use.

### Electrochemical measurements

Electrochemical behavior of the NiONPs-CPE was examined in 0.1 mol L<sup>-1</sup> KCl solution containing 1 mmol L<sup>-1</sup> Fe(CN)<sub>6</sub><sup>3-/4-</sup> by CV and EIS methods and compared to those of UCPE. The cyclic voltammograms (CVs) were recorded in potential range between -0.2 V and 0.6 V at different scan rates. The EIS studies were performed at the frequency range between 100 kHz and 0.01 Hz with 5 mV perturbation at open-circuit potential ( $E_{\text{OCP}}$ ).

Electrocatalytic activity of the NiONPs-CPE was investigated in the absence and presence of 1 mmol L<sup>-1</sup> H<sub>2</sub>O<sub>2</sub> in 0.1 mol L<sup>-1</sup> phosphate buffer solution (PBS, pH 7.0) by CV and chronoamperometry methods and compared to that of UCPE. The CVs were recorded in potential range between -0.3 V and 1.0 V at a 50 mVs<sup>-1</sup> scan rate. The amperometric measurements were carried out at +0.4 V vs. Ag/AgCl.

Amperometric determination of glucose was performed by the oxidation of enzymatically produced H<sub>2</sub>O<sub>2</sub> at +0.4 V vs. Ag/AgCl in an aqueous solution containing 0.1 mol L<sup>-1</sup> PBS (pH 7.0). First, the GOD enzyme electrode was equilibrated in PBS at +0.4 V until a constant current ( $i_1$ ), which was called as background current, was obtained. Then, aliquots of glucose solution were added to electrochemical cell. The steady-state current response ( $i_2$ ) to the addition of glucose was recorded and the current difference ( $\Delta i = i_2 - i_1$ ) was determined with the change between the steady-state current and the background current. A calibration curve of  $\Delta i$ -glucose concentration was plotted.

## Results and discussion

### The effect of the quantity of NiONPs on the performance of the modified CPE electrode

The quantity of NiONPs to be added into carbon paste would affect the performance of the modified electrode, so the NiONPs-CPEs were prepared at the various mass ratio of NiONPs/graphite (1:5, 1:4, 1:3, and 1:2). In order to examine the electron transfer features of the prepared NiONPs-CPEs were used Fe(CN)<sub>6</sub><sup>3-/4-</sup> as redox probe. The cyclic voltammetry of ferricyanide is a useful tool to monitor the electrochemical behavior of the modified electrodes (Liu et al. 2006). Figure 1(A) shows the CVs of the NiONPs-CPEs in the solution of redox couple. A pair of redox waves of Fe(CN)<sub>6</sub><sup>3-/4-</sup> was observed between 0.0 V and 0.4 V at all electrodes. Both cathodic peak current and anodic peak current generally increase with increasing quantity of NiONPs. It can be said that NiONPs act as an electrocatalyst. The electrochemical data obtained from Figure 1(A) are presented Table I. While the anodic and cathodic peak currents at the ratios of 1:3 and 1:4 are the highest, the peak-to-peak potential separation ( $\Delta E_p$ ) between the cathodic and anodic waves at the ratio of 1:4 is the lowest. The NiONPs-CPEs prepared with the ratios of 1:3 and 1:4 exhibited the comparatively lowest  $\Delta E_p$  and highest peak currents, indicating fast electron transfer kinetics. To determine the most appropriate combination of electrode, the amperometric responses of the three NiONPs-CPEs to the successive addition of H<sub>2</sub>O<sub>2</sub> were recorded at an applied potential of +0.4 V. Figure 1(B) displays the relationships between  $\Delta i$  and H<sub>2</sub>O<sub>2</sub> concentration

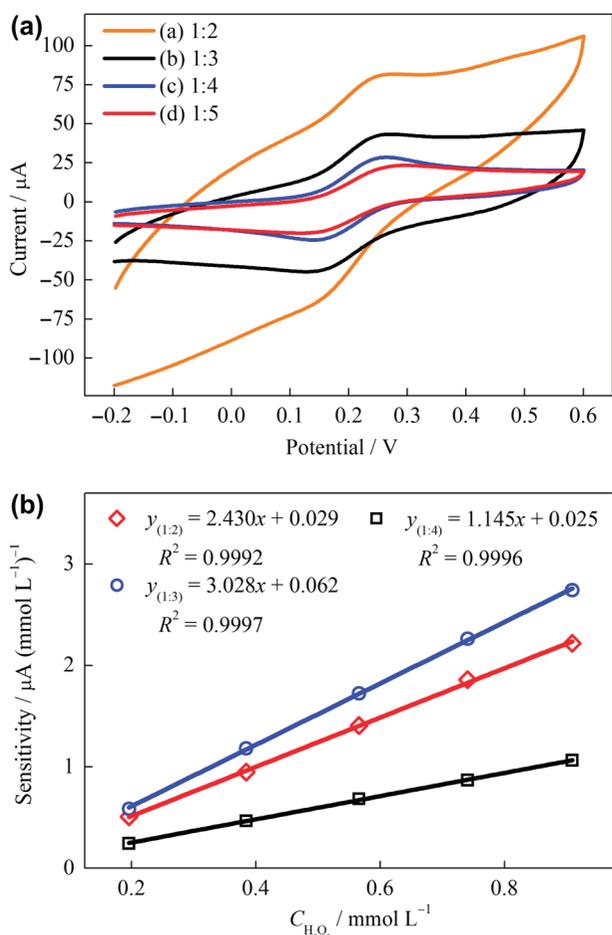


Figure 1. (A) CVs of  $1 \text{ mmol L}^{-1} \text{ Fe(CN)}_6^{3-/4-}$  on the various NiONPs-CPEs with the mass ratios of NiONPs to graphite powder were 1:2 (a), 1:3 (b), 1:4 (c), and 1:5 (d) (Scan rate:  $50 \text{ mV s}^{-1}$ ), (B) the plot of the peak current vs.  $\text{H}_2\text{O}_2$  concentration for the various NiONPs-CPEs in  $0.1 \text{ M PBS (pH 7.0)}$  at  $0.4 \text{ V}$  vs. Ag/AgCl.

( $0.20\text{--}0.91 \text{ mmol L}^{-1}$ ). The highest sensitivity was obtained when the NiONPs-CPE at the ratio of 1:3 was used. The mass ratio of NiONPs and graphite powder are significant for electrochemical detection based on the oxidation of  $\text{H}_2\text{O}_2$ . Furthermore, a large amount of nanoparticles incorporated into the carbon paste can lead to unfavorable mechanical properties of CPE and increase the resistance of the electrode (Li et al. 2011).

### Electrochemical properties of unmodified and modified electrodes

The effect of NiONPs on the electron transfer features of carbon paste electrode was investigated in KCl solution containing  $\text{Fe(CN)}_6^{3-/4-}$ . The CVs of the UCPE and

Table I. Comparisons of electrochemical characteristics of the NiONPs-CPEs were prepared at the various mass ratio of NiONPs/graphite toward the redox probe of  $\text{Fe(CN)}_6^{3-/4-}$  (scan rate =  $50 \text{ mVs}^{-1}$ ).

The mass ratio of NiONPs to graphite	$E_{pa}/\text{V}$	$E_{pc}/\text{V}$	$\Delta E_p/\text{V}$	$I_{pa}/\mu\text{A}$	$I_{pc}/\mu\text{A}$
1:2	0.245	0.147	0.098	16.44	14.66
1:3	0.254	0.147	0.107	20.51	21.47
1:4	0.256	0.149	0.107	20.13	20.35
1:5	0.281	0.139	0.142	15.86	15.46

NiONPs-CPE were shown in Figure 2(A). The redox peaks of  $\text{Fe(CN)}_6^{3-/4-}$  were clearly observed for UCPE and NiONPs-CPE. The current intensity of the redox waves for the NiONPs-CPE extremely was increased when compared to that of the UCPE. In addition, while the peak-to-peak separation ( $E_{pa} - E_{pc} = \Delta E_p$ ) of  $\text{Fe(CN)}_6^{3-/4-}$  waves for the UCPE was  $320 \text{ mV}$ , the  $\Delta E_p$  for the NiONPs-CPE was  $100 \text{ mV}$ . These results denoted that NiONPs provide the increasing electron transfer at solution/electrode interface due to their electrocatalytic and conductivity properties and/or the extended electroactive surface area of the modified electrode (Yao and Shiu 2007, Zhang et al. 2009, Zheng and Zhang 2007). The effect of scan rate on the voltammetric response of the NiONPs-CPE was shown in Figure 2(B). As the scan rate increases, the peak currents increase while the redox peak potentials shift slightly. The anodic and cathodic peak currents (Figure 2(B), inset) exhibit a linear dependence on the square root of the scan rate, indicating the mass transfer phenomenon at the electrochemical probe and the electrode interface is mainly a diffusion-controlled process (Du et al. 2007, Song et al. 2010).

The EIS is an effective technique to investigate the interface properties of modified electrodes (Wang 2006b). The electrochemical impedance spectra (Nyquist curves) of the UCPE and NiONPs-CPE are shown in Figure 3. The Nyquist

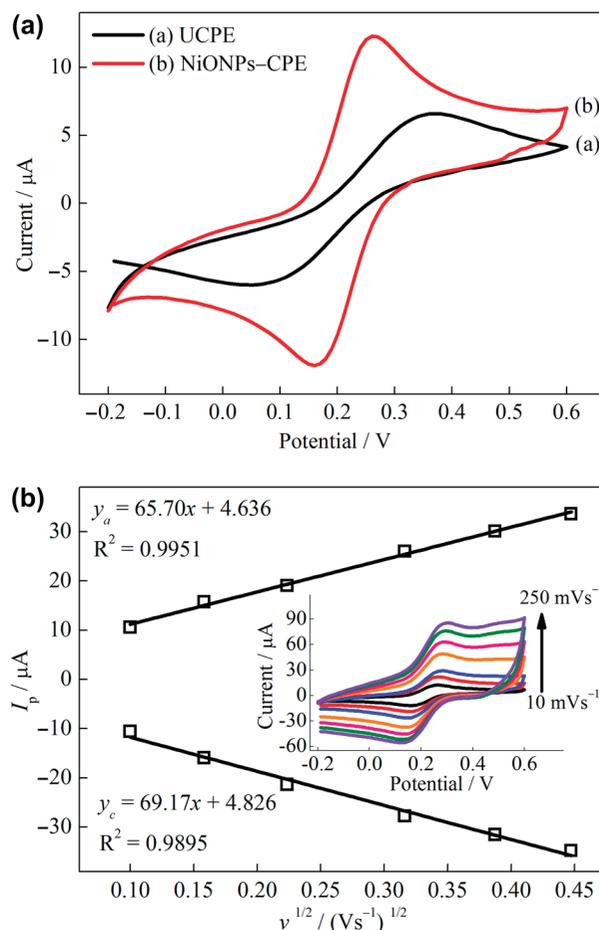


Figure 2. (A) CVs of the UCPE (a) and NiONPs-CPE (b) at scan rate  $10 \text{ mVs}^{-1}$ , (B) the plot of the peak current vs. the square root of the scan rate for NiONPs-CPE (inset: CVs at different scan rates) in  $0.1 \text{ mol L}^{-1} \text{ KCl}$  solution containing  $1 \text{ mmol L}^{-1} \text{ Fe(CN)}_6^{3-/4-}$ .

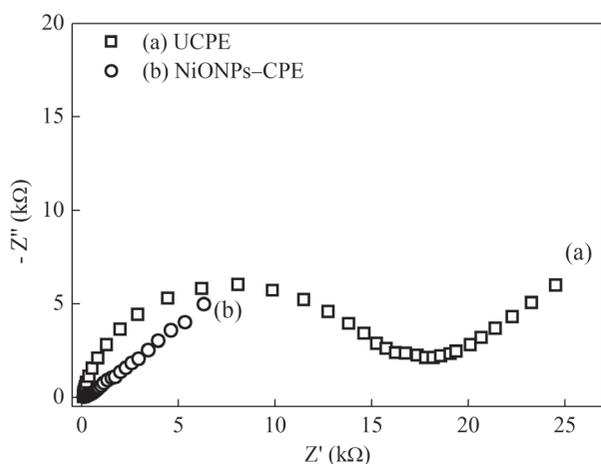


Figure 3. The Nyquist curves of the (a) UCPE and (b) NiONPs-CPE in  $0.1 \text{ mmol L}^{-1}$  KCl solution containing  $1 \text{ mmol L}^{-1}$   $\text{Fe}(\text{CN})_6^{3-/4-}$ .

curves include a semicircle portion at high frequencies and a linear portion at low frequencies. The semicircle portion and the linear portion correspond to the electron transfer limited process and the diffusion process, respectively. The semicircle diameter is equal to electron transfer resistance ( $R_{\text{et}}$ ) value (Wang 2006b). The UCPE has a large semicircle with a large electron transfer resistance ( $R_{\text{et}}$ ) of approximately  $16 \text{ k}\Omega$  (Figure 3a). Whereas, the NiONPs-CPE exhibits almost a straight line, indicating that NiONPs provide fast electron transfer kinetics for the  $\text{Fe}(\text{CN})_6^{3-/4-}$  redox system. This is attributable to the good electron transfer ability of NiONPs which can facilitate the electrochemical oxidation of  $\text{H}_2\text{O}_2$  on the electrode surface (Li et al. 2011).

### Electrocatalytic activity of unmodified and NiONPs-modified electrodes

The electrocatalytic characteristic of the NiONPs-CPE electrode was examined using cyclic voltammetry (CV) and chronoamperometry (CA) methods for  $\text{H}_2\text{O}_2$  oxidation and compared to that of UCPE. Figure 4(A) exhibits the CVs of NiONPs-CPE and UCPE in the absence and presence  $\text{H}_2\text{O}_2$  ( $1 \text{ mmol L}^{-1}$ ). As seen in Figure 4(A) (a) and (c), while a small background current was obtained for the UCPE in PBS in the potential range from  $-0.3$  to  $1.0 \text{ V}$ , a significant increase of current response was obtained for the NiONPs-CPE. It is indicated that the NiONPs have excellent electrochemical properties, which could be ascribed to the large surface area, high electrical conductivity and enhanced electron transfer of NiONPs (Noorbakhsh and Salimi 2011, Zhang et al. 2009) Furthermore, a pair of redox couples due to the  $\text{Ni}^{2+}/\text{Ni}^{3+}$  redox process was observed between  $0.4$  and  $0.9 \text{ V}$  at the NiONPs-CPE (Adekunle and Ozoemena 2008, Adekunle and Ozoemena 2010). Upon the addition of  $1 \text{ mmol L}^{-1}$   $\text{H}_2\text{O}_2$ , a little increase of current intensity at the UCPE was observed after about  $+0.6 \text{ V}$ , indicating the oxidation of  $\text{H}_2\text{O}_2$  (Figure 4(A) (b)). On the other hand, a remarkably increase of current intensity at the NiONPs-CPE was shown after about  $+0.2 \text{ V}$ , which could be attributed to the electrocatalytic activity of NiONPs (Figure 4(A) (d)) (Zheng and Zhang 2007). The developed behavior for  $\text{H}_2\text{O}_2$  oxidation on the surface of NiONPs-CPE may be due to an increase in the rate of electron transfer from  $\text{H}_2\text{O}_2$  to the electrode surface

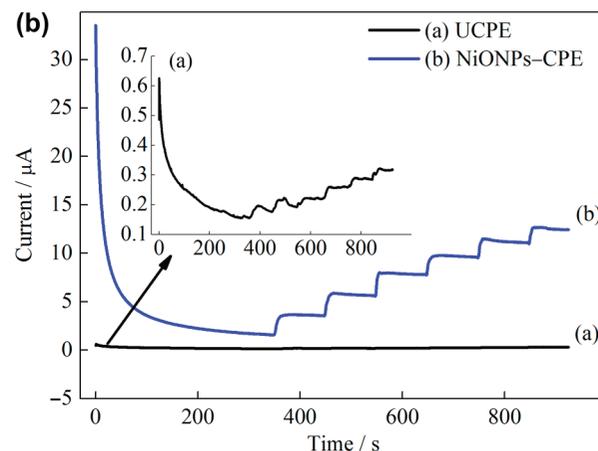
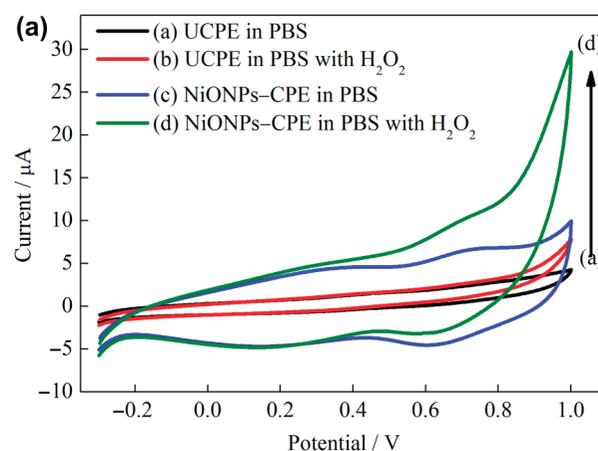


Figure 4. (A) CVs of the UCPE and NiONPs-CPE in the absence and presence of  $1 \text{ mmol L}^{-1}$   $\text{H}_2\text{O}_2$  (scan rate =  $50 \text{ mVs}^{-1}$ ) in  $0.1 \text{ mmol L}^{-1}$  PBS (pH 7.0). (B) Amperometric responses of the UCPE and NiONPs-CPE for the oxidation of  $\text{H}_2\text{O}_2$  in  $0.1 \text{ M}$  PBS (pH 7.0) at  $0.4 \text{ V}$  vs. Ag/AgCl.

incorporated the NiONPs. The NiONPs-CPE also provided a noticeable decrease in the overpotential for  $\text{H}_2\text{O}_2$  oxidation. This indicates that it allows amperometric detection of  $\text{H}_2\text{O}_2$  at low potential.

In order to investigate further the electrocatalytic activity of the NiONPs-modified electrode, amperometric measurements were performed at the NiONPs-CPE. Figure 4(B) shows the amperometric responses to  $\text{H}_2\text{O}_2$  of the NiONPs-CPE and UCPE to the successive addition of  $\text{H}_2\text{O}_2$  at an applied potential of  $0.4 \text{ V}$ . As seen in Figure 4(B), the current responses to  $\text{H}_2\text{O}_2$  of the NiONPs-CPE were higher than that of the UCPE. The NiONPs-modified electrode developed the current density corresponds to the oxidation of  $\text{H}_2\text{O}_2$ , indicating the electrocatalytic activity of NiONPs. Consequently, the NiONPs can be used for construction of sensors including amperometric detection of  $\text{H}_2\text{O}_2$  at low potential.

### Effect of working potential on the response of biosensor

The dependency of the biosensor response on the working potential was carried out in  $0.1 \text{ mol L}^{-1}$  PBS, pH 7.0. Figure 5 shows the effect of working potential on the amperometric response of the biosensor in the potential ranging from  $0.3$  to  $0.6 \text{ V}$  using varying  $\text{H}_2\text{O}_2$  concentrations from  $0.2 \text{ mmol L}^{-1}$  to  $1.38 \text{ mmol L}^{-1}$ . As expected, the current response of the biosensor increased with the applied potential up to a

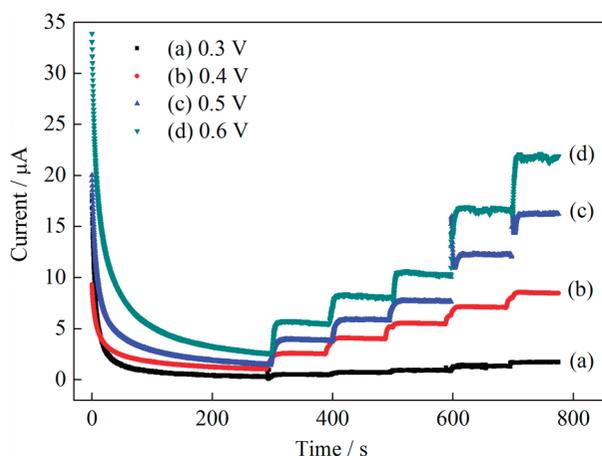


Figure 5. The effect of working potential on the amperometric response of  $\text{H}_2\text{O}_2$  in  $0.1 \text{ mmol L}^{-1}$  PBS (pH 7.0).

value of 0.6 V vs. Ag/AgCl. However, the interference effect of the most common species, present in many body fluids, is remarkably considered at high potentials. Accordingly, a potential of a 0.4 V, where the interference species are not oxidizable and resulting in less interference, was used in all measurements (Xu et al. 2004).

#### Effect of enzyme unit, pH and temperature on the response of biosensor

The sensitivities of the biosensor were determined at different GOD units (5–15 U). The effect of enzyme loading on biosensor sensitivity was investigated by measuring the different GOD concentrations (5–15 U)-loaded biosensors responses to sequential glucose increments from 0.02 to 0.20  $\text{mmol L}^{-1}$ . The other parameters were kept constant at +0.4 V. As expected from the increased biocatalytic activity of the electrode, the sensitivity was enhanced with the increasing of the amount of enzyme in the carbon paste matrix. Although the maximum sensitivity was obtained at 15 U GOD loading, 12 U was selected due to negligible increment (Supplementary Figure 1 to be found online at <http://www.informahealthcare.com/doi/abs/10.3109/21691401.2013.808649>).

The influence of the pH was checked in  $0.1 \text{ mol L}^{-1}$  PBSs with glucose increments from 0.02 to 0.20  $\text{mmol L}^{-1}$  in the

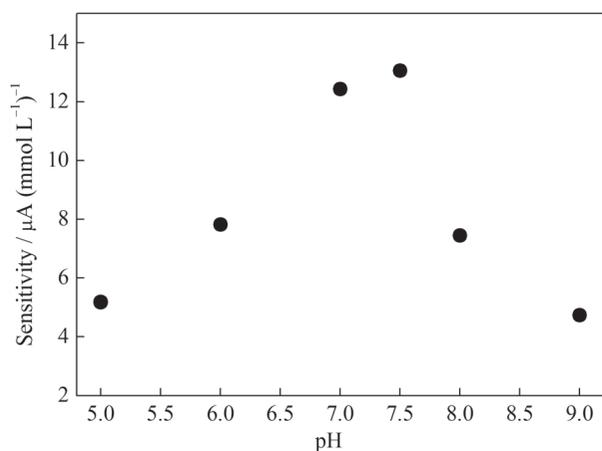


Figure 6. The sensitivity of the GOD-NiONPs-CPE as a function of buffer pH in  $0.1 \text{ mmol L}^{-1}$  PBS at 0.4 V vs. Ag/AgCl.

range of pH 5.0–9.0. As can be seen from Figure 6, the sensitivity of the biosensor was increased with the pH value was increased up to pH 7.5. The immobilized GOD presented maximum activity at pH 7.5. This is in agreement with that reported in literatures (Pan et al. 2004, Li and Lin 2007). However, pH 7.0 was selected in the glucose detection due to its linear working range was relatively wider than that of others (not shown).

The increment in the response of sensors could be affected by temperature increases due to enhanced enzymatic activity, increased membrane permeability of various participating species (Vaddiraju et al. 2011). Investigation of the temperature dependency on the response of the glucose biosensor was obtained by measuring the electrode responses to sequential glucose increments from 0.02 to 0.20  $\text{mmol L}^{-1}$  maintained at the desired temperature (10–50°C). The sensitivity of the biosensor increased with increasing of temperature until 45°C, and then, it declined slowly (Supplementary Figure 2 to be found online at <http://www.informahealthcare.com/doi/abs/10.3109/21691401.2013.808649>). The decline in the response of sensors could be due to denaturation of GOD. However, for practical convenience, the temperature of 25°C was selected as a working temperature for all further experiments.

#### Effect of glucose concentration on the response of biosensor

The effect of the glucose concentration on the amperometric response was studied using varying concentration of glucose ( $1 \times 10^{-5}$ – $1 \times 10^{-1} \text{ mol L}^{-1}$ ). As seen in Figure 7(A), two linear working ranges were obtained. The inset in Figure 7(A) (a) shows a good linear response in the range of  $1.9 \times 10^{-3}$ – $9.1 \times 10^{-3} \text{ mmol L}^{-1}$  glucose with a correlation coefficient of 0.966 and the regression equation is  $I (\mu\text{A}) = 367.13C_{\text{glucose}} + 13.05$ . Also, a good linear response in the range of 1.31–15.00  $\text{mmol L}^{-1}$  glucose with a correlation coefficient of 0.966 was obtained (Figure 7(A) (b)) and the regression equation is  $I (\mu\text{A}) = 0.15C_{\text{glucose}} + 18.80$ . Forzani et al. have reported that the normal clinical range for glucose in blood is between 3.5 and 6.1  $\text{mmol L}^{-1}$  (Forzani et al. 2004). The linear range of the GOD-NiONPs-CPE covers the normal serum glucose level.

The limit of detection (LOD) is calculated to be  $1.1 \times 10^{-4} \text{ mmol L}^{-1}$  ( $S/N=3$ ). LOD values confirmed the sensitivity of the GOD-NiONPs-CPE which is calculated using following equation:

$$\text{LOD} = 3s/m;$$

where  $s$  is the standard deviation of the current (five runs) of the lowest concentration of the linearity range and  $m$  is the slope of the related calibration curve.

$K_m$  gives an indication of the enzyme-substrate kinetics for the glucose biosensor. Kinetic parameters such as  $K_m$  and  $I_{\text{max}}$  were obtained from the Hanes-Woolf plot by plotting “substrate concentration/current response vs. substrate concentration” (Figure 7B).  $K_m$  and  $I_{\text{max}}$  were calculated as 0.081 mM and 20.79  $\mu\text{A}$ , respectively, using following equation (Garrett and Grisham 2009).

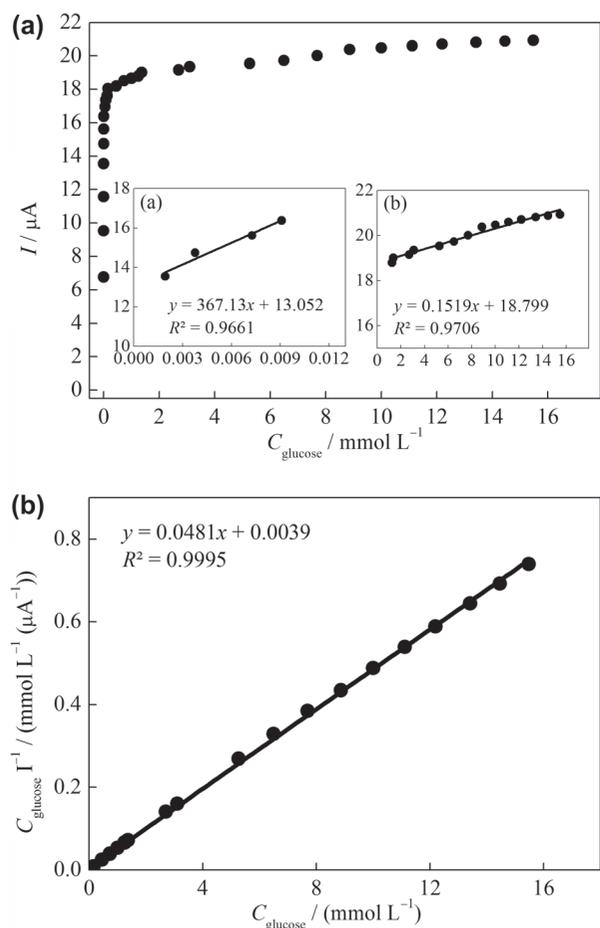


Figure 7. (A) Calibration curve of the GOD-NiONPs-CPE for glucose in 0.1 M PBS (pH 7.0) at 0.4 V vs. Ag/AgCl. (inset: linear working ranges plots), (B) Substrate concentration/current response vs. substrate concentration plot (Hanes-Woolf plot).

$$\frac{[S]}{I} = \left( \frac{1}{I_{\max}} \right) [S] + \frac{K_m}{I_{\max}}$$

It is desired to obtain the highest  $i_{\max}$  and lowest  $K_m$  for construction of enzyme electrodes (Wang 2006a). The  $K_m$  in the present work is lower than GOD immobilized on NiO hollow nanospheres ( $7.76 \text{ mmol L}^{-1}$ ) (Li et al. 2008), NiO-doped ZnO nanorods ( $7.4 \text{ mmol L}^{-1}$ ) (Chu et al. 2012), and electrodeposited NiO nanoparticles ( $2.7 \text{ mmol L}^{-1}$ ) (Salimi et al. 2007b).

### Repeatability, reproducibility and storage stability

The repeatability of the enzyme electrode was evaluated using the sensitivities of same glucose biosensor for eight times consecutively at +0.4 V. The relative standard deviation of the obtained sensitivities was found as 6.58%. This result was the indication of good stability.

The reproducibility of the biosensor was also carried out by measuring the biosensor responses to sequential glucose increments from 0.02 to 0.20  $\text{mmol L}^{-1}$  in the same day. Seven different biosensors, fabricated independently, had a relative standard deviation of 10.66%. These results indicate that the repeatability and reproducibility of the biosensor were highly satisfactory and obtained biosensor can be used in many analyses.

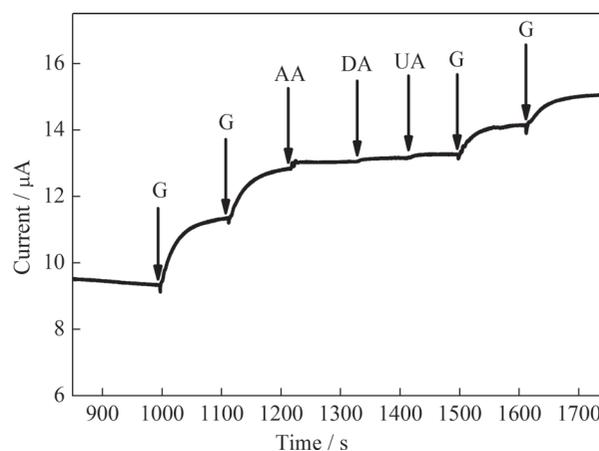


Figure 8. Amperometric responses of the glucose biosensor upon addition of AA ( $0.1 \text{ mmol L}^{-1}$ ), DA ( $0.1 \text{ mmol L}^{-1}$ ), and UA ( $0.1 \text{ mmol L}^{-1}$ ) in 0.1 M PBS (pH 7.0) at 0.4 V vs. Ag/AgCl.

The long-term stability is one of the important features required for the satisfactory application of a biosensor. The stability of the glucose biosensor was performed by measuring the amperometric responses to sequential glucose increments from 0.02 to 0.20  $\text{mmol L}^{-1}$ . The long-term stability of the GOD-NiONPs-CPE was determined by performing activity assays within 20 days. The electrode was stored in dry atmosphere at  $+4^\circ\text{C}$  when not in use. An activity loss of 10% was observed on the 20th day. The long-term stability of the GOD-NiONPs-CPE is much better compared to the stability of another carbon material electrode with a 26.5% loss of response after 22 days (Tang et al. 2004). Good long-term stability can be attributed to the mild immobilization procedure and a beneficial environment for preventing enzyme leakage. Immobilization of GOD in carbon paste matrix provides a biocompatible microenvironment around the enzyme. The developed biosensor achieved 95% of the steady-state current within about 5 s which is quite satisfactory for biosensor response. Such a fast response may reflect the increased electron transfer in the NiONPs-modified carbon paste.

### Effect of interferences on the response of biosensor

The real sample contains other common species that could potentially interfere with the glucose detection. In the light of low concentrations of the ascorbic acid, dopamine, and uric acid in the blood, the interference needs to be seriously considered because they could be easily oxidized at a relatively high over potential, and a lower working potential helps to avoid the electrochemical reaction of the interferences.

Table II. Effects of interferences on glucose response.

Interfering species	Concentration <sup>a</sup> ( $\text{mmol L}^{-1}$ )	Interference <sup>b</sup> (%)
L-ascorbic acid	0.1	1.56
Dopamine	0.1	1.54
Uric acid	0.1	0.76

<sup>a</sup>The electrochemical cell concentration of interfering species.

<sup>b</sup>% interference =  $(i_c - i_0)/i_0 \times 100$ ;  $i_0$ : The amperometric response of buffer solution containing  $0.1 \text{ mmol L}^{-1}$  glucose solution and  $i_c$ : The amperometric response of buffer solution containing  $0.1 \text{ mmol L}^{-1}$  glucose and interfering species with above concentration.

Table III. The determination of glucose levels in human serum samples.

Serum samples	Glucose concentration (mmol L <sup>-1</sup> )	
	Provided by the hospital	Determined by proposed biosensor*
1	3.61	3.50 ± 0.6
2	4.61	4.93 ± 0.9
3	5.00	5.14 ± 0.4
4	7.44	7.53 ± 1.0

\*Average of three measurements.

The selectivity of proposed biosensor was performed in the presence of common interfering species such as ascorbic acid (AA, 0.1 mmol L<sup>-1</sup>), dopamine (DA, 0.1 mmol L<sup>-1</sup>), and uric acid (UA, 0.1 mmol L<sup>-1</sup>) in oxygenated PBS considering their physiological levels (Figure 8). The percentages of interference were determined in terms of the difference in percentage of amperometric response between the buffer solution containing glucose ( $i_0$ ) and buffer solution containing interfering species and glucose ( $i_x$ ). Thus, the equation,  $(i_x - i_0)/i_0 \times 100$  was used. The biosensor shows quick response for each addition of 0.1 mmol L<sup>-1</sup> glucose (G). However, as can be seen from Table II no noteworthy response was observed for the addition of ascorbic acid, dopamine, and uric acid into buffer solution containing glucose at +0.4 V (vs. Ag/AgCl). Although the ascorbic acid, dopamine, and uric acid caused a small increase in the current, these results can be overseen for the determination of glucose in serum samples due to maximum concentrations of L-ascorbic acid and uric acid in human body is about 0.1 and 0.05 mmol L<sup>-1</sup>, respectively (Xiao et al. 2000, Aydođdu et al.). Thus, the developed biosensor is highly selective for glucose and can be applied for glucose determination in human serum samples under physiological conditions.

### Determination of glucose in human serum

The determination of glucose in the serum sample was investigated on the GOD-NiONPs-CPE utilizing a standard addition method. Human serum samples were provided by a local hospital, which were first analyzed with Beckmann Coulter Olympus AU2700 analyzer which is based on hexokinase method. Then, serum samples of 10  $\mu$ L were successively added into 5 mL phosphate buffered (0.1 mmol L<sup>-1</sup> pH 7.0), and the response of the proposed biosensor was obtained at 0.4 V. The accuracy of glucose biosensor was controlled by  $t$ -test. The  $t$  value was found as 0.408 at 95% confidence level, for  $t_{\text{critic}}$  4.30. As illustrated in Table III, a good correlation between the two methods results allows us to ascertain the practical applicability of the developed biosensor.

### Conclusion

In the present study, a novel NiONPs-modified carbon paste enzyme electrode was developed for the determination of glucose. The NiONPs and GOD incorporated carbon paste electrode acted as a biosensor. The performance of CPE was dramatically improved due to the excellent electrocatalytic activity of NiONPs. The NiONPs-modified glucose biosensor showed good characteristics such as a wide determination

range with a low detection limit (0.11  $\mu$ mol L<sup>-1</sup>), high sensitivity, and stability. Furthermore, it has good reproducibility and applicability to human serum analysis and provides the determination of glucose at a low potential (+0.4 V). Kinetic parameters of the presented glucose biosensor  $K_m$  and  $I_{\text{max}}$  were calculated as 0.081 mmol L<sup>-1</sup> and 20.79  $\mu$ A, respectively. The easy fabrication and excellent electrocatalytic capability of the prepared glucose biosensor can be expected to find practical applications to fabricate novel electrochemical biosensors.

### Declaration of interest

The authors report no declarations of interest. The authors alone are responsible for the content and writing of the paper

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## Supplementary material available online

Supplementary Figure 1 and 2